Naming Oxy-ions

	The Element at the center of the ion	Chg the element prefers	Chg of the ion			
3/4 ions	P	-3	-3	4 oxygens	PO_4^{3-}	Phosphate ion
				3 oxygens	PO_3^{3-}	Phosphite ion
	S	-2	-2	4 oxygens	SO_4^{2-}	Sulfate ion
				3 oxygens	SO_3^{2-}	Sulfite ion
2/3 ions	C	-4/+4	-2	3 oxygens	CO_3^{2-}	Carbonate ion
				2 oxygens	CO_2^{2-}	Carbonite ion
	N	-3	-1	3 oxygens	NO_3^{1-}	Nitrate ion
				2 oxygens	NO_2^{1-}	Nitrite ion
	Cl, Br, I	-1	-1	4 oxygens	BrO_4^{1-}	*Per*bromate ion
				3 oxygens	BrO_3^{1-}	Bromate ion
				2 oxygens	BrO_2^{1-}	Bromite ion
				1 oxygen	BrO_1^{1-}	*Hypo*bromite ion

Think of a party, and remember: I "ate more". The ate ion has more oxygens.

NH_4^{1+}	Ammonium ion
OH^{1-}	Hydroxide ion
OAc^{1-} ($CH_3CO_2^{1-}$)	Acetate Ion
CN^{1-}	Cyanide Ion

- All 20 oxy-ions have the same charge as their central atom normally prefers – except carbon and nitrogen
- P and S oxy-ions have 3 or 4 oxygens.
- C,N,Cl,Br, and I oxy-ions have 2 or 3 oxygens.
- The "ite" ion is always the one with less oxygens.
- The "ate" ion is always the one with more oxygens.
- The 3 halogens can super size: 4 oxygens = perchlorate ion
- The 3 halogens can also have a really small ion: 1 oxygen = hypochlorite ion
- There are four misc polyatomic ions you should also know:

  NH_4^{1+}, OH^{1-}, OAc^{1-}, and CN^{1-}.