

Nomenclature

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“What’s in a name?” — William Shakespeare

Communicating in chemistry is facilitated by a chemical vocabulary to name compounds. Naming compounds is called nomenclature from the Latin *nomen* (*name*) and *calere* (*to call*). There are millions of known chemical substances, so naming them all would require hundreds of pages. In this lab we will focus our discussion to the nomenclature of inorganic compounds- compounds that do not generally contain carbon.

Common Names

Before their compositions were known, many compounds were given informal, *common names*, many of which are still in use among chemists. Table 1 lists the common names, formulas, and chemical names (systematic names) of some familiar compounds.

Table 1

Common Name	Chemical Name	Formula
baking soda	sodium hydrogen carbonate sodium bicarbonate	NaHCO ₃
bleach (solid)	sodium perborate	NaBO ₃
Borax	sodium tetraborate decahydrate	Na ₂ B ₄ O ₇ ·10 H ₂ O
brimstone	sulfur	S
cream of tartar	potassium hydrogen tartrate	KHC ₄ H ₄ O ₆
Epsom salt	magnesium sulfate heptahydrate	MgSO ₄ ·7 H ₂ O
Freon	dichlorodifluoromethane	CF ₂ Cl ₂
grain alcohol	ethanol	C ₂ H ₅ OH
graphite	carbon	C
gypsum	calcium sulfate dihydrate	CaSO ₄ ·2 H ₂ O
laughing gas	dinitrogen oxide	N ₂ O
lime	calcium oxide	CaO
limestone	calcium carbonate	CaCO ₃
lye	sodium hydroxide	NaOH
marble	calcium carbonate	CaCO ₃
MEK	ethyl methyl ketone	CH ₃ COC ₂ H ₅
milk of magnesia	magnesium hydroxide	Mg(OH) ₂
muriatic acid	hydrochloric acid	HCl
oil of vitriol	sulfuric acid	H ₂ SO ₄

plaster of paris	calcium sulfate ½ hydrate	CaSO ₄ ·½ H ₂ O
potash	potassium carbonate	K ₂ CO ₃
iron pyrite (fool's gold)	iron disulfide	FeS ₂
quartz	silicon dioxide	SiO ₂
quicksilver	mercury	Hg
rubbing alcohol	isopropyl alcohol	(CH ₃) ₂ CHOH
sal ammoniac	ammonium chloride	NH ₄ Cl
salt	sodium chloride	NaCl
saltpeter	potassium nitrate	KNO ₃
slaked lime	calcium hydroxide	Ca(OH) ₂
sugar	sucrose	C ₁₂ H ₂₂ O ₁₁
TSP trisodium phosphate	sodium phosphate	Na ₃ PO ₄
wood alcohol	methyl alcohol	CH ₃ OH

Systematic Names

Chemists prefer systematic names that precisely identify the chemical composition of compounds. The systematic name (also called the IUPAC name) is the “official” name of a chemical, as determined by the naming rules published by IUPAC (International Union of Pure and Applied Chemistry). A systematic name reveals which elements are present and, in some cases, the arrangement of atoms in a compound. In the remainder of this lab, we will focus on the systemic naming of compounds using IUPAC rules.

I - Nomenclature of Cations

The name of a monatomic cation is the same as the name of the element forming it, with the addition of the word “ion”, as in “sodium ion” for Na⁺. For representative elements, the magnitude of the charge is not specified since it is directly related to the group number in which the element is located. When an element can form more than one kind of cation, such as Cu⁺ and Cu²⁺ from copper (a transition metal element), we use the charge of the cation, written as a Roman numeral in parenthesis immediately following the name of the element. For example, Cu⁺ is a copper(I) ion and Cu²⁺ is a copper (II) ion. Table 2 shows some more examples:

Table 2

Atom	Name	Ion	Name
K	Potassium	K ⁺	Potassium ion
Ba	Barium	Ba ²⁺	Barium ion
Fe	Iron	Fe ²⁺	Iron(II) ion
Fe	Iron	Fe ³⁺	Iron(III) ion

II – Nomenclature of Anions

Anions are named differently from cations. To name an anion consisting of one element, use the stem of the parent element name and change the ending to “*ide*”. As an example, the F^- ion is named by using the stem *fluor-* from fluorine and adding *-ide* to form fluoride ion. The charge of the monoatomic anion is not included in its name but is figured out as: group number – 8. For example, the formula for fluoride ion is F^- , the minus charge was figured out from 7 (group number) – 8 = 1-

Table 3 provides examples of various non-metal elements that form anions along with the stem of the non-metals name.

Table 3

Symbol	Element	Stem	Anion name	Anion Symbol
Br	Bromine	Brom	Bromide	Br^-
Cl	Chlorine	Chlor	Chloride	Cl^-
F	Fluorine	Fluor	Fluoride	F^-
H	Hydrogen	Hydr	Hydride	H^-
I	Iodine	Iod	Iodide	I^-
N	Nitrogen	Nitr	Nitride	N^{3-}
O	Oxygen	Ox	Oxide	O^{2-}
P	Phosphorous	Phosph	Phosphide	P^{3-}
S	Sulfur	Sulf	Sulfide	S^{2-}
C	Carbon	Carb	Carbide	C_2^{2-}
Se	Selenium	Selen	Selenide	Se^{2-}
As	Arsenic	Arsen	Arsenide	As^{3-}
Te	Tellurium	Tellur	Telluride	Te^{2-}

III - Nomenclature of Binary Ionic Compounds

Binary compounds are those formed between two elements. If one of the elements is a metal and the other a nonmetal, the binary compound is usually made up of ions, a cation and an anion and it is called a “binary ionic compound.” Nomenclature for binary ionic compounds falls into two categories; 1) binary ionic compounds containing a metal forming only one type of cation, Type I and 2) binary ionic compounds containing a metal that can form two or more types of cations, Type II.

Nomenclature Rules for Metals Forming Only One Type of Cation – Type I

Metals that form only one type of cation tend to be the alkali metals (Group IA, form +1 ions) and the alkaline earths (Group IIA, form +2 ions), along with aluminum (+3 ion), zinc (+2 ion), cadmium (+2 ion) and silver (+1 ion). Follow the steps below when naming binary ionic compounds containing metals with fixed charges:

1. The order for names in a binary compound is first the cation (metal), then the anion (non-metal).
2. For the cation, write the name of the parent atom.
3. Write the name of the anion by first writing the stem for the anion and then adding the suffix “-ide.”

Example: Name the compound **NaCl**

1. The compound is composed of Na, a metal and Cl, a non-metal. Sodium is in group IA of the periodic table and forms only one type of cation with a +1 charge, Na^+ . We name the cation “*sodium*.”
2. The name of the non-metal is chlorine. We modify chlorine to its stem “chlor” and add “ide” to form the name of the anion “*chloride*.”
3. The name of the compound is *sodium chloride*.

Example: Name the compound **CaS**

1. The compound is composed of Ca, a metal and S, a non-metal. Calcium is in group IIA of the periodic table and forms only one type of cation with a +2 charge, Ca^+ . We name the cation “*calcium*.”
2. The name of the non-metal is sulfur. We modify sulfur to its stem “sulf” and add “ide” to form the name of the anion “*sulfide*.”
3. The name of the compound is “*calcium sulfide*.”

Practice Problems – Given the Formula, Write the Name

Write the correct name for the following Type I binary ionic compounds. Remember that all elements involved in this problem set have only one charge. That includes both the cation and the anion in the formula.

- 1) MgS _____
- 2) KBr _____
- 3) Ba_3N_2 _____
- 4) Al_2O_3 _____
- 5) NaI _____

- 6) SrF₂ _____
- 7) Li₂S _____
- 8) RaCl₂ _____
- 9) CaO _____
- 10) AlP _____
- 11) K₂S _____
- 12) LiBr _____
- 13) Sr₃P₂ _____
- 14) BaCl₂ _____
- 15) NaBr _____
- 16) MgF₂ _____
- 17) Na₂O _____
- 18) SrS _____
- 19) BN _____
- 20) AlN _____

Now let's practice writing formulas for Type I binary ionic compounds if given the name. To do this, remember the following rules:

1. The order in a formula is first the cation, then the anion.
2. You must know the charges associated with each cation and anion.
3. The sum of the positive charge and the sum of the negative charges **MUST** add up to zero.
4. You may not adjust the charges of the cations or anions to get a total charge of zero.
5. You may adjust the subscripts to get a total charge of zero.

Example 1: Write the formula from the following name: **sodium bromide**

Step #1 - Write down the symbol and charge of the first word. Result = Na⁺

Step #2 - Write down the symbol and charge of the second word. Result = Br⁻

Step #3 - Use the minimum number of cations and anions needed to make the sum of all charges in the formula equal zero. In this case, only one Na⁺ and one Br⁻ are required.

The resulting formula is NaBr.

Example 2: Write the formula from the following name: **aluminum chloride**

Step #1 - Write down the symbol and charge of the first word. Result = Al³⁺

Step #2 - Write down the symbol and charge of the second word. Result = Cl⁻

Step #3 - Use the minimum number of cations and anions needed to make the sum of all charges in the formula equal zero. In this case, only one Al³⁺ is required, but three Cl⁻ are required.

Why? Answer - Three negative one charges are required because there is one positive three charge. Only in this way can the total charge of the formula be zero.

The resulting formula is AlCl_3 .

Example 3: Write the formula from the following name (using least common multiple technique): **aluminum sulfide**.

Step #1 - Write Al^{3+} and S^{2-} right next to each other.

Step #2 - Select the least common multiple between the two charges (ignoring the signs). In this case it is 6. This is the amount of positive charge and negative charge you need for a correct formula. Thus you need sufficient aluminum atoms to make a +6 charge and you need sufficient sulfur atoms to make a -6 charge.

The resulting formula requires two Al and three S, so the correct formula is Al_2S_3 .

Practice Problems – Given the Name, Write the Formula

Write the correct formula for the following Type I binary ionic compounds. Remember that all elements involved in this problem set have only one charge. That includes both the cation and the anion in the formula.

- 1) magnesium oxide _____
- 2) lithium bromide _____
- 3) calcium nitride _____
- 4) aluminum sulfide _____
- 5) potassium iodide _____
- 6) strontium chloride _____
- 7) sodium sulfide _____
- 8) radium bromide _____
- 9) magnesium sulfide _____
- 10) aluminum nitride _____
- 11) cesium sulfide _____
- 12) potassium chloride _____
- 13) strontium phosphide _____
- 14) barium iodide _____
- 15) sodium fluoride _____
- 16) calcium bromide _____
- 17) beryllium oxide _____
- 18) strontium sulfide _____
- 19) boron fluoride _____

20) aluminum phosphide

Nomenclature Rules for Metals That Can Form Two or More Types of Cations – Type II

The metals in the center of the periodic table (the transition metals) often form more than one type of cation. For example, lead can form Pb^{+2} and Pb^{+4} ions, and copper can form Cu^+ and Cu^{2+} ions. This can be confusing when you are naming compounds. For example lead chloride could be PbCl_2 or PbCl_4 . To resolve this problem the IUPAC devised the **Stock System** to name these compounds. In the stock system, when a compound contains a metal that can form more than one type of cation, the charge on the cation of the metal is designated by a Roman numeral placed in parenthesis immediately following the name of the metal. The anion (non-metal) of the compound is named in the usual manner for binary compounds.

As an example the metal iron forms two common ions, Fe^{2+} and Fe^{3+} . The first is called the iron(II) ion, and the second is the iron(III) ion. The Roman numeral immediately following the name of the metal indicates the charge on the ion. Thus, FeCl_2 is iron(II) chloride, while FeCl_3 is iron(III) chloride.

Example #1: Write the name for: **SnCl_2**

Step #1 - the first part of the name is the unchanged name of the first element (the metal cation) in the formula. In this example, it would be tin.

Step #2 – since tin forms either a +2 or +4 cation the result from step one will be followed by a Roman numeral, either (II) or (IV). The charge on Cl is -1 and since there are 2 Cl atoms in the formula the total charge equals -2. Since the total charge of the formula unit must be zero (it is neutral) the charge on Sn must be +2. Therefore we use tin(II) for the name of the metal. Notice that there is no space between the name and the parenthesis.

Step #3 - the anion is named in the usual manner of stem plus "ide." In this example the name of the anion is derived in the following manner; take the stem of chlorine and add ide to it: chlor + ide = chloride.

The correct name of the example is tin(II) chloride.

Example #2: Write the name for: **Fe_2O_3**

Step #1 - the first part of the name is the unchanged name of the first element in the formula. In this example, it would be iron.

Step #2 – since iron forms either a +2 or +3 cation the result from step one will be followed by a Roman numeral, either (II) or (III). The charge on O is -2 and since there are 3 O atoms in the formula the total charge equals -6. Since the total charge of the formula unit must be zero (it is neutral) the charge on each Fe atom must be +3 ($2 \times +3 = +6$). Therefore we use iron(III) for the name of the metal.

Step #3 - the anion is named in the usual manner of stem plus "ide." In this example the name of the anion is derived in the following manner; take the stem of oxygen and add ide to it: ox + ide = oxide.

The correct name of the example is iron(III) oxide.

Practice Problems – Given the Formula, Write the Name

Write the correct name for the following Type II binary ionic compounds. Remember that metals involved in this problem can form two or more cations.

- 1) CuS _____
- 2) PbBr₄ _____
- 3) Pb₃N₂ _____
- 4) Fe₂O₃ _____
- 5) FeI₂ _____
- 6) Sn₃P₄ _____
- 7) Cu₂S _____
- 8) SnCl₂ _____
- 9) HgO _____
- 10) Hg₂F₂ _____
- 11) CuCl₂ _____
- 12) CuBr _____
- 13) PbO _____
- 14) Fe₂S₃ _____
- 15) PbCl₂ _____
- 16) SnO _____
- 17) Cu₂O _____
- 18) PbO₂ _____
- 19) FeO _____
- 20) SnO₂ _____

Now let's practice writing formulas for Type II binary ionic compounds if given the name. Here are a few examples;

Example #1 - Write the formula for: **copper(II) chloride**

Step #1 - the first word tells you the symbol of the cation and the Roman numeral will tell you the charge on the cation. In this case we write Cu²⁺.

Step #2 - the anion symbol and charge comes from the second name. In this case, chloride means Cl⁻.

Step #3 - remembering the rule that a formula must have zero total charge, you write the formula CuCl₂.

Example #2 - Write the formula for: **iron(III) sulfide**

Step #1 - the symbol of the cation is Fe and it has a +3 charge (Roman numeral is 3) and so we write Fe⁺³.

Step #2 - Sulfide (the anion) means S^{2-} .

Step #3 - since a formula must have zero total charge, you write the formula Fe_2S_3 .

Example #3 - Write the formula for: **tin(IV) phosphide**

Step #1 -First symbol is Sn from the name tin and the Roman numeral IV gives +4 as tin's charge. We write Sn^{4+} .

Step #2 - Phosphide gives P^{3-} .

This compound's formula is Sn_3P_4 .

Practice Problems – Given the Formula, Write the Name

Write the correct formula for the following Type II binary ionic compounds.

1) iron(II) chloride _____

2) copper(I) sulfide _____

3) lead(IV) iodide _____

4) tin(II) fluoride _____

5) chromium(II) nitride _____

6) tin(II) oxide _____

7) chromium(III) oxide _____

8) gold(I) iodide _____

9) manganese(II) nitride _____

10) cobalt(III) phosphide _____

11) iron(III) chloride _____

12) copper(II) sulfide _____

13) lead(II) bromide _____

14) tin(IV) iodide _____

15) mercury(II) fluoride _____

16) tin(IV) oxide _____

17) manganese(III) chloride _____

18) chromium(II) nitride _____

19) gold(III) oxide _____

20) cobalt(II) phosphide _____

IV - Nomenclature Rules for Binary Molecular Compounds Containing Two Nonmetals

If the two elements in a binary compound are both nonmetals instead of a metal and a nonmetal, the compound is a molecular compound. Examples of binary molecular compounds include CO, NO, BF₃ and OCl₂. Nomenclature for these compounds follows the rules below:

1. In a compound formed between two nonmetals, the element that occurs first in this series is written and named first:

Si, B, P, H, C, S, I, Br, N, Cl, O, F

2. Write the name of the first element and use a Greek prefix to indicate the number of atoms of each element (except that mono is not used).

Prefix	Number
Mono	1
Di	2
Tri	3
Tetra	4
Penta	5
Hexa	6
Hepta	7
Octa	8
Nona	9
Deca	10

When the prefix ends with an “a” or “o” and the second element begins with a vowel, the a or o is often dropped.

3. Write the stem of the second element and add the suffix “ide.” Use a prefix to indicate the number of atoms for the second element.

Example #1 - write the name for **NO₂**.

Step #1 - The first element in the formula is nitrogen. Since there is only one nitrogen atom present we do not need to add a prefix, i.e., we do not write “mono” nitrogen just nitrogen.

Step #2 – The second element in the formula is oxygen. Since two oxygen atoms are present in the formula we add the prefix “di.”

The correct name for the compound is nitrogen dioxide.

Example #2 - write the name for **N₂O**.

Step #1 - The first element in the formula is nitrogen. Since there are two nitrogen atoms present we add the prefix “di” giving dinitrogen.

Step #2 – The second element in the formula is oxygen. Since one oxygen atom is present in the formula we add the prefix “mono.”

Step #3 – We drop one “o” and write monoxide rather than monooxide.

The correct name for the compound is dinitrogen monoxide.

Example #3 - write the name for N_2O_5 .

Step #1 - the first element is nitrogen and there are two. This part of the name will be "dinitrogen."

Step #2 - the second element is oxygen, so "oxide" is used. Since there are five, the prefix "penta" is used.

Step #3 – We drop the “a” from penta and write pentoxide.

The correct name for the compound is dinitrogen pentoxide.

Practice Problems – Given the Formula, Write the Name

Write the correct name for the following binary molecular compounds.

Write the correct name for:

- 1) As_4O_{10} _____
- 2) BrO_3 _____
- 3) BN _____
- 4) N_2O_3 _____
- 5) NI_3 _____
- 6) SF_6 _____
- 7) XeF_4 _____
- 8) PCl_3 _____
- 9) CO _____
- 10) PCl_5 _____
- 11) P_2O_5 _____
- 12) S_2Cl_2 _____
- 13) ICl_2 _____
- 14) SO_2 _____
- 15) P_4O_{10} _____
- 16) UF_6 _____
- 17) OF_2 _____

18) ClO₂ _____

19) SiO₂ _____

20) BF₃ _____

Now let's practice writing formulas for binary molecular compounds if given the name. Here are a few examples;

Example #1 - write the formula for **dinitrogen trioxide**.

Example #2 - write the name for **carbon monoxide**.

Step #1 - the first name will tell you the first element in the formula. In the first example above, it would be N and in the second, C.

If there is a prefix on the name, this gives the subscript on the element. In the first example above, the "di-" tells you there are two nitrogen's. Absence of a prefix, as in the second example, says there is only one of that element involved.

Step #2 - the anion name tells you the element; oxide means oxygen. Once again, the prefix will tell you how many of the element are involved. "Tri-" means three and "mono-" means one.

The correct formulas of the two examples are N₂O₃ and CO.

Note that "monoxide" is written rather than "oxide" when there is one atom of the second element involved. Note also that when one element of the first atom is involved, no "mono-" is used. Monocarbon monoxide is just as wrong as carbon oxide.

Example #3 - write the formula for **bromine pentafluoride**.

Step #1 - the first symbol is Br and its subscript will be a one, which is understood to be present.

Step #2 - the second element is fluorine, so F is used. The prefix "penta-" indicates a subscript of 5.

The formula of this compound is BrF₅.

Example #4 - write the formula for **diphosphorous pentoxide**.

Step #1 - the first symbol is P and the subscript is 2.

Step #2 - pentoxide says five oxygen's are involved.

The formula of this compound is P₂O₅.

Practice Problems – Given the Name, Write the Formula

Write the correct formula for the following binary molecular compounds.

1) chlorine monoxide _____

2) oxygen difluoride _____

3) boron phosphide _____

4) dinitrogen monoxide _____

- 5) nitrogen trifluoride _____
- 6) sulfur tetrachloride _____
- 7) xenon trioxide _____
- 8) carbon dioxide _____
- 9) diphosphorous pentoxide _____
- 10) phosphorous trichloride _____
- 11) sulfur dioxide _____
- 12) bromine pentafluoride _____
- 13) disulfur dichloride _____
- 14) boron trifluoride _____
- 15) tetraarsenic decoxide _____
- 16) silicon tetrachloride _____
- 17) krypton difluoride _____
- 18) chlorine monoxide _____
- 19) silicon dioxide _____
- 20) boron trichloride _____

V – Nomenclature for Naming Compounds Containing Polyatomic Ions

A polyatomic ion is an ion that contains two or more elements joined together by covalent bonds. Table 4 provides the names, formulas and charges of some common polyatomic ions;

	The Element at the center of the ion	Chg the element prefers	Chg of the ion			
3/4 ions	P	-3	-3	4 oxygens	PO ₄ ³⁻	Phosphate ion
				3 oxygens	PO ₃ ³⁻	Phosphite ion
	S	-2	-2	4 oxygens	SO ₄ ²⁻	Sulfate ion
				3 oxygens	SO ₃ ²⁻	Sulfite ion
2/3 ions	C	-4/+4	-2	3 oxygens	CO ₃ ²⁻	Carbonate ion
				2 oxygens	CO ₂ ²⁻	Carbonite ion
	N	-3	-1	3 oxygens	NO ₃ ¹⁻	Nitrate ion
				2 oxygens	NO ₂ ¹⁻	Nitrite ion
	Cl, Br, I	-1	-1	4 oxygens	BrO ₄ ¹⁻	Perbromate ion
				3 oxygens	BrO ₃ ¹⁻	Bromate ion
2 oxygens				BrO ₂ ¹⁻	Bromite ion	
1 oxygen				BrO ₁ ¹⁻	Hypobromite ion	
	NH ₄ ¹⁺				Ammonium ion	
	OH ¹⁻				Hydroxide ion	
	OAc ¹⁻ (CH ₃ CO ₂ ¹⁻)				Acetate ion	
	CN ¹⁻				Cyanide ion	

Think of a party, and remember: I "ate more". The ate ion has more oxygens.

Ionic compounds containing polyatomic ions generally consist of one or more cations combined with a negative polyatomic ion. These compounds are not binary compounds. They contain three or more elements, as opposed to only two in a binary compound.

When naming these compounds, it is important that you learn to recognize the presence of a polyatomic ion in a formula and know their charges. Also, in these compounds, the cations used will be a mix of fixed charges (Type I) and variable charges (Type II). You must know which are which.

As an example, consider the formula NaHCO_3 . The formula consists of two parts, a cation, Na^+ and a polyatomic ion, HCO_3^- . When naming compounds containing polyatomic ions we use the following rules;

1. Write the name of the cation
2. Write the name of the polyatomic anion

The correct name for the formula NaHCO_3 is therefore sodium bicarbonate.

When more than one unit of the corresponding polyatomic ion is required, parenthesis are used to enclose the ion with the subscript going outside the parenthesis. For example, the very first formula used is $\text{Fe}(\text{NO}_3)_2$. This means that two NO_3^- are involved in the compound. Here are some examples;

Example #1 - write the name for **$\text{Fe}(\text{NO}_3)_2$** Step #1 - decide if the cation is one showing variable charge. If so, a Roman numeral will be needed. In this case, iron does show variable charge.

If a variable charge cation is involved, you must determine the Roman numeral involved. You do this by computing the total charge contributed by the polyatomic ion. In this case, NO_3^- has a minus one charge and there are two of them, making a total of minus 2.

Therefore, the iron must be a positive two, in order to keep the total charge of the formula at zero.

Step #2 - determine the name of the polyatomic ion. Nitrate is the name of NO_3^- .

The correct name is iron(II) nitrate.

Example #2 - write the name for **NaOH**

Step #1 - the cation, Na^+ , does not show a variable charge, so no Roman numeral is needed. The name is sodium.

Step#2 - OH^- is recognized as the hydroxide ion.

The name of this compound is sodium hydroxide.

Example #3 - write the name for **KMnO_4**

Step #1 - the cation, K^+ , does not show a variable charge, so no Roman numeral is needed. The name is potassium.

Step#2 - MnO_4^- is recognized as the permanganate ion.

The name of this compound is potassium permanganate.

Example #4 - write the name for Cu_2SO_4

Step #1 - decide if the cation is one showing variable charge. If so, a Roman numeral will be needed. In this case, copper does show variable charge.

If a variable charge cation is involved, you must determine the Roman numeral involved. You do this by computing the total charge contributed by the polyatomic ion. In this case, SO_4^{2-} has a minus two charge and there is only one, making a total of minus 2.

Therefore, the copper must be a positive one. Why? Well, there must be a positive two to go with the negative two in order to make zero. Since the formula shows two copper atoms involved, each must be a plus one charge.

Step #2 - determine the name of the polyatomic ion. Sulfate is the name of SO_4^{2-} .

The correct name is copper(I) sulfate.

Example #5 - write the name for $\text{Ca}(\text{ClO}_3)_2$

The first part of the name comes from the first element's name: calcium. You also determine that it is not a cation of variable charge.

The second part of the name comes from the name of the polyatomic ion: chlorate.

This compound is named calcium chlorate.

Practice Problems – Given the Formula, Write the Name

The cations in this first set are all of fixed oxidation state, so no Roman numerals are needed.

Write the correct name for:

- 1) AlPO_4 _____
- 2) KNO_2 _____
- 3) NaHCO_3 _____
- 4) CaCO_3 _____
- 5) $\text{Mg}(\text{OH})_2$ _____
- 6) Na_2CrO_4 _____
- 7) $\text{Ba}(\text{CN})_2$ _____
- 8) K_2SO_4 _____
- 9) NaH_2PO_4 _____
- 10) NH_4NO_3 _____

The cations in this set are all of variable oxidation state, so Roman numerals are needed.

Write the correct name for:

- 11) $\text{Sn}(\text{NO}_3)_2$ _____
12) FePO_4 _____
13) Cu_2SO_4 _____
14) $\text{Ni}(\text{C}_2\text{H}_3\text{O}_2)_2$ _____
15) HgCO_3 _____
16) $\text{Pb}(\text{OH})_4$ _____
18) $\text{Cu}(\text{ClO}_3)_2$ _____
19) FeSO_4 _____
20) $\text{Hg}_2(\text{ClO}_4)_2$ _____

These formulas mix the use of the two types of cations.

Write the correct name for:

- 21) KClO_3 _____
22) SnSO_4 _____
24) $\text{Pb}(\text{NO}_3)_2$ _____
25) $\text{Mg}_3(\text{PO}_4)_2$ _____
26) CuH_2PO_4 _____
27) CaHPO_4 _____
28) $\text{Fe}(\text{HCO}_3)_3$ _____
29) Na_2CO_3 _____
30) MnSO_4 _____

Now let's practice writing names for compounds containing polyatomic ions if given the formula. Here are a few examples;

Example #1 - write the formula for **copper(II) chlorate**

Step #1 - the first word tells you the symbol of the cation. In this case it is Cu.

Step #2 - the Roman numeral **WILL** tell you the charge on the cation. In this case it is a positive two.

Step #3 - the polyatomic formula and charge comes from the second name. In this case, chlorate means ClO_3^- .

Step #4 - remembering the rule that a formula must have zero total charge, you write the formula $\text{Cu}(\text{ClO}_3)_2$.

Example #2 - write the formula for **silver cyanide**

Step #1 - the first word tells you the symbol of the cation. In this case it is Ag^+ .

Step #2 - silver has a constant charge of +1, it is not a cation with variable charge.

Step #3 - the polyatomic formula and charge comes from the second name. In this case, cyanide means CN^- .

Step #4 - remembering the rule that a formula must have zero total charge, you write the formula AgCN .

Example #3 - write the formula for **sodium phosphate**

Step #1 - the cation, sodium, is Na^+ , and it does not show a variable charge.

Step#2 - phosphate is PO_4^{3-} .

The formula of this compound is Na_3PO_4 . Notice that no parenthesis are required since only one polyatomic is used.

Example #4 - write the name for **barium carbonate**

Step #1 - the cation, barium, does not show a variable charge and its symbol is Ba^{2+} .

Step#2 - carbonate is CO_3^{2-} .

The formula of this compound is BaCO_3 .

The cations in this first set are all of fixed oxidation state, so no Roman numerals are needed.

Practice Problems – Given the Name, Write the Formula

The cations in this first set are all of fixed oxidation state, so no Roman numerals are needed.

Write the correct formula for:

- 1) silver carbonate
- 2) potassium hydrogen phosphate
- 3) aluminum hydroxide
- 4) sodium hydrogen carbonate
- 5) calcium acetate
- 6) potassium permanganate
- 7) calcium perchlorate
- 8) lithium carbonate
- 9) magnesium hydrogen sulfite
- 10) sodium hypochlorite

The cations in the following set are all of variable oxidation state, so Roman numerals are needed.

Write the correct formula for:

- 11) tin(IV) chlorite _____
- 12) mercury(II) phosphate _____
- 13) tin(II) carbonate _____
- 14) mercurous acetate _____
- 15) lead(II) chromate _____
- 16) copper(I) sulfite _____
- 17) stannous dichromate _____
- 18) iron(III) nitrate _____
- 19) ferric sulfate _____
- 20) ferrous hydroxide _____

These formulas mix the use of the two types of cations.

Write the correct formula for:

- 21) potassium perchlorate _____
- 22) lead(IV) hydrogen phosphate _____
- 23) aluminum sulfate _____
- 24) iron(II) bicarbonate _____
- 25) barium iodate _____
- 26) tin(II) hydrogen sulfide _____
- 27) magnesium dihydrogen phosphate _____
- 28) silver phosphate _____
- 29) cobalt(III) nitrite _____
- 30) ammonium sulfate _____
- 31) ammonium nitrate _____